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13) Using the equation from problem #12, determine the mass of aluminum acetate that can be made if I do this reaction with 125 grams of acetic acid and 275 grams of aluminum hydroxide. 14) What is the limiting reagent in problem #13? 15) How much of the excess reagent will be left over in problem #13 after the reaction is complete?

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Sample Problem #2. Write the chemical equation for the following reaction: Hydrochloric acid and solid sodium hydrogen carbonate react to produce aqueous sodium chloride, water, and carbon dioxide. Hint: Acids will always be aqueous unless otherwise stated.  $\text{HCl (aq)} + \text{NaHCO}_3\text{(s)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)} + \text{CO}_2\text{(g)}$

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#### Chapter 11: Chemical Reactions

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Balancing Equations: Answers to Practice Problems 1. Balanced equations. (Coefficients equal to one (1) do not need to be shown in your answers).

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Thermochemistry is the study of how heat changes during phase changes and chemical reactions. Types of Energy + Quantifying Energy (ability to do work or produce heat) Calories Joules Law of Conservation of Energy – 1 st Law of Thermodynamics ex) Potential Energy stored in gasoline bonds ? Kinetic energy in ...

Chemical Reactions Guided Practice Problem CHEMICAL REACTIONS (pages 330-339). This section GUIDED PRACTICE PROBLEM 2 (page 324). 2. possibly a change of state. Essential Questions and Answers. 1. What are the types for energy? Chemical Energy is energy caused by chemical reactions. Quiz: Chemical Reactions. 1/8. During chemical reactions, atoms.