
Chemistry Chapter 12 Stoichiometry

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The calculation of quantities in chemical reactions is a subject of chemistry. Mole ratio. A conversion factor derived from coefficients of a balanced chemical equation interpreted in terms of moles. Chemistry Chapter 12 Stoichiometry Stoichiometry 379 CHAPTER 12 Assessment 36 Chapter 12 stoichiometry test b answer key. a. Two formula units KClO_3 decompose to form two formula units KCl and three molecules O_2 . b Chapter 12 stoichiometry test b

answer key. Four molecules NH_3 react with six molecules NO to form five mol-
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Stoichiometry CHAPTER 12 What You'll Learn You will write mole ratios from balanced chemical equations. You will

calculate the number of moles and the mass of a reactant or product when given the number of moles or the mass of another reactant or product. You will identify the limiting reactant in a chemical reaction. You will determine the
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solution are
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same manner as
we discussed in
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11. Instead of
masses,
however, we use
volumes of
solutions of
known
concentration to
determine the
number of moles
of reactants.
Whether we are
dealing with
volumes of
solutions of
reactants or

masses of
reactants, the
coefficients in the
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equation tell us
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