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The problem was finally solved by dissolving KF in liquid HF to give a

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The cell voltage is given by: $2.2 \text{ [Cu]} 0.0257 \text{ V dilute} \ln 2 \text{ [Cu]} \text{ concentrated} + + = ^\circ - E E 0.0257 \text{ V } 0.080 0 \ln 2 1.2 = - = 0.035 \dots$

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For each solution, calculate $[\text{H}^+]$ or $[\text{OH}^-]$. State whether the ...

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