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16. An isotope has an atomic number 51 and a mass number 123.

*Chemistry Matter Change Chapter 3 Study Guide Answer*

*Key*

1. Atoms are divisible into smaller, subatomic particles. 2. A given element can have atoms with different masses, called isotopes. The formation of water according to the equation.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . shows that 2 molecules (made of 4 atoms) of hydrogen and 1 molecule (made of 2 atoms) of oxygen produce 2 molecules of water.

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radius and ionic radius.  
Circle the letter of the  
choice that best completes  
the  
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Change Chapter 12 Section  
3 Study ...  
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guide. matter can be  
changed from one form into  
another, mixtures can be  
separated or made, and  
pure substances can be  
decomposed, but the total  
amount of mass remains  
constant. If two elements

form more than one  
compound between them,  
then the ratios of the  
masses of the second  
element which combine with  
a fixed mass of the first  
element will be ratios of  
small whole numbers.  
Chemistry Matter And  
Change Chapter 3 Study  
Guide Answer Key  
1. all matter is composed of  
atoms, 2. atoms of a given  
element are identical ,  
atoms of different elements  
differ from each other, 3.  
atoms cannot be  
subdivided, created, or  
destroyed, 4. atoms  
combine in simple whole-  
number ratios to form

compounds, 5. in chemical  
reactions elements are  
combined, separated, or  
rearranged.  
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Compounds - An

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Introduction to Chemistry. Chapter 3 – Chemical Compounds 25 To get a review of the most important topics in the chapter, fill in the blanks in the Key Ideas section. Work all of the selected problems at the end of the chapter, and check your answers with the solutions provided in this chapter of the study guide.

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