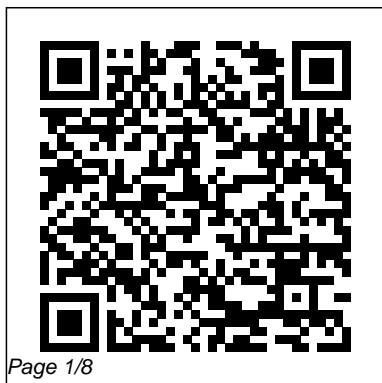

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(53) the idea of arranging the elements in a table according to their chemical and physical properties is attributed to

_____.
Chapter 6 Oxidation-Reduction Reactions
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Periodic Table and Periodic Law Section 6.1 Development of the Modern Periodic Table In your textbook, reads about the history of the periodic table's development. Use each of the terms below just once to complete the passage. nine eight accepted octaves elements protons atomic mass properties periodic law atomic number Henry Moseley Dmitri Mendeleev The table below was

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Chemistry Chapter 6 Study Guide. arranged elements according to atomic mass and used the arrangement to predict the properties of missing elements. elements that are characterized by the filling of p orbitals are classified as ____.

which subatomic particle plays the greatest part in determining the properties of an element.

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able to... Define chemistry.
Describe the difference
between mass and weight.
Describe the difference

between chemical and
physical change. Describe the
difference between accuracy
and precision. Describe the
five states of matter, and give
an example of each

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1.octaves. 2.eight. 3.nine.
4.accepted. 5. Dmitri
Mendeleev. 6.atomic mass.
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 6 Unit 2 Name Date Class In your
 textbook, read about water's
 polarity. Label the diagram. Use
 these choices: covalent bond.

hydrogen bond. slightly negative end. slightly positive end. 1. slightly negative end 2. slightly positive end 3. hydrogen bond 4. covalent bond In your textbook, read about mixtures with water.

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CHAPTER 6 Section 1:

Atoms, Elements, and

Compounds nucleus proton

nucleus I eve In your textbook,

read about the structure of

atoms. Label the diagram of an

atom. Use these choices:

electron energy level neutron

13 In your textbook, read about

elements, compounds, and

chemical bonds. If the

statement is true, write true.