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LTHS Answers

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Concentration and Time (Last Page) - Unit 8.2 Practice Problems Key

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Calculate the average atomic masses. Round all

answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80% ^{127}I and 20% ^{129}I . Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that have the following

HW Keys - Roosevelt High School AP Chemistry 2017-18

ChemQuest 36 Name: _____ Date: _____ Hour: _____

Information: Double Replacement Reactions

In the previous ChemQuest, you learned about single replacement reactions in which a single atom replaces an ion from another reactant. Study what happens in the following reactions. They are called double replacement reactions.

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44. Critical Thinking Questions. 5. In each of the Bohr diagrams in Figure 1, the first energy level only has two electrons drawn in it. Why is this. No more than 2 electrons can fit in the first energy level of any atom. 6. What is the maximum number of electrons that the second energy level can have How many electrons can the 3rd energy level have

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ChemQuest 15 Name: _____ Date: _____ Hour: _____

_____ Information: Relating Electron

Configurations to the Periodic Table In this section you will see how the periodic table serves as a road map for writing electron configurations. Get your periodic table out and get ready. Remember that a row on the periodic table

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(Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) 12 moles. For every mole of propane that combusts 3 moles of CO_2 are produced, so the number of moles of CO_2 that can be produced when 4 moles of propane combusts = $4(3) = 12$.

ChemQuest 15 - Yav Science

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ChemQuest 35 Intro, to chemical Reactions

Date: _____ Hour: _____ Formula Review 1. Write the formulas for the following compounds: 2.

Write the names for the following ...