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AHS Chemistry Resource Site - Unit 6 - Rates & Equilibrium

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Collision Theory Model: Collision Theory In the picture below, the baseball bat represents Reactant A and the baseball represents Reactant B A reaction will only be successful if the batter hits a homerun. If the batter does not hit a homerun, the reaction will be considered a failure. Now, read

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Collision Theory - Impact for a Chemical Reaction Why? The collision theory states that a chemical reaction can only occur between particles when they collide (hit each other). The collision between reactant particles is necessary but not sufficient for a reaction to take place. The collisions also have to be effective.

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1) Collision Theory POGIL . 2) Factors the Affect Rate . 3) Potential Energy Diagrams . 4) Dynamic Equilibrium POGIL . 5) LeChatelier's Principle POGIL . 6) LeChatelier's Principle . 7) Equilibrium Constants . 8) ICE Problems . 9) Ka and Kb Problems

Collision Theory - Impact for a Chemical Reaction

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taking place

Based on your answer to Key Question 5, complete the following statement: Collision theory states that a reaction is most likely to occur if&mlr; If the reactants collide with the correct amount of energy and correct orientation. Exercise 1. Hydrogen gas and iodine vapor combine to form hydrogen iodide gas, as shown in the equation $\text{H}_2 + \text{I}_2 \rightarrow 2 \text{HI}$.