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collision theory is that, in order to react, molecules must collide. This fundamental rule guides any analysis of an ordinary reaction mechanism. Collision Theory Pogil Solution The collision theory states that a chemical reaction can only occur between particles when they collide (hit each other). The collision between reactant particles is necessary but not

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Collision Theory - Impact for a Chemical Reaction Why? The collision theory states that a chemical reaction can only occur between particles when they collide (hit each other). The collision between reactant particles is necessary but not sufficient for a reaction to take place. The collisions also have to be effective.

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Based on your answer to Key Question 5, complete the following statement: Collision theory states that a reaction is most likely to occur if&mdash; If the reactants collide with the correct amount of energy and correct orientation. Exercise 1. Hydrogen gas and iodine vapor combine to form hydrogen iodide gas, as shown in the equation  $H_2 + I_2 \rightarrow 2 HI$ .

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Collision Theory Model: Collision Theory In the picture below, the baseball bat represents Reactant A and the baseball represents Reactant B A reaction will only be successful if the batter hits a homerun. If the batter does not hit a homerun, the reaction will be considered a failure. Now, read

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