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**The Electrical  
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Aqueous Solutions  
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*The Electrical Conductivity of Non-Aqueous Solutions*  
CRC Press

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The High-frequency

Conductivity of Aqueous Solutions of Colloidal Electrolytes The Electrical Conductivity of Aqueous Solutions The Electrical Conductivity of Aqueous Solutions of Strong Electrolytes at High Frequencies The Electrical Conductivity of Aqueous Solutions Excerpt from The Electrical Conductivity of

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Aqueous Solutions: A Report Presented by Arthur A. Noyes Upon a Series of Experimental Investigations Executed by A. A. Noyes, W. D. Coolidge, A. C. Melcher, H. C. Cooper, Yogoro Kato, R. B. Sosman, G. W. Eastman, C. W. Kanolt, and W. Bottger

The investigation to be described in the following series of articles was undertaken for the purpose of studying through a wide range of temperature, extending from 18 to the critical temperature and above, the electrical conductivity of aqueous solutions and such other physical and chemical properties of them as are related to it or can be determined through measurements of it. Aside from its direct physical significance, it is well known that the electrical conductivity of solutions is a property of fundamental importance in connection with the ionic theory; for it gives the simplest and most direct measure of the ionization of substances, upon which their chemical behavior in solution depends. A full investigation of this property at all



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temperatures would therefore furnish a comprehensive knowledge of the chemical equilibrium of dissolved substances in water; and if supplemented by determinations of the solubility of solid salts, which determinations can also be made by measuring the conductance of their saturated solutions, a fairly complete basis for the development of the

chemistry of aqueous solutions of electrolytes would be obtained. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at [www.forgottenbooks.com](http://www.forgottenbooks.com) This book is a reproduction of an important historical work. Forgotten Books uses state-of-the-art technology to digitally reconstruct the work, preserving the original format

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The Electrical  
Conductivity of  
Aqueous Solutions of  
Uranyl Perchlorate

Hardpress Publishing

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*The Electrical  
Conductivity of  
Aqueous Solutions*  
Palala Press  
Properties of  
Aqueous Solutions  
of Electrolytes is  
a handbook that  
systematizes the

information on physico-chemical parameters of multicomponent aqueous electrolyte solutions. This important data collection will be invaluable for developing new methods for more efficient chemical technologies, choosing optimal solutions for more effective methods of using raw materials and

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energy resources, and other such activities. This edition, the first available in English, has been substantially revised and augmented. Many new tables have been added because of a significantly larger list of electrolytes and their properties (electrical conductivity, boiling and freezing points, pressure of saturated vapors, activity and diffusion coefficients). The book is divided into two sections. The first section provides tables that list the properties of binary aqueous solutions of electrolytes, while the second section deals with the methods for calculating their properties in multicomponent systems. All values are given in PSI units or fractional and multiple units. Metrological characteristics of the experimental methods used for the determination of physico-chemical parameters are indicated as a relative error and those of the computational

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methods as a relative error or a root-mean square deviation.

**The Electrical Conductivity of Aqueous Solutions of Strong Electrolytes at High Frequencies**

Excerpt from The Electrical Conductivity of Non-Aqueous Solutions Van't Hoff formulated his modern theory of solutions from the results of Pfeffer's classic experiments on the osmotic pressure of a few substances in

aqueous solutions, and observed, however, from the work of de Vries<sup>3</sup> who pointed out that equimolecular quantities of dissolved substances produce the same osmotic pressure. Van't Hoff<sup>4</sup> stated that the condition of a substance in the gaseous state is analogous to the condition in which a substance exists in dilute solutions wherein the osmotic pressure of the dissolved substance corresponds to the pressure of the gaseous particles. It was

observed, however, that many substances did not give normal values for the osmotic pressure, the rise of the boiling, or the lowering of the freezing point; and thus that the laws of gases would not hold for solutions of these substances. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at [www.forgottenbooks.com](http://www.forgottenbooks.com) This book is a reproduction of an important historical

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### **Properties of Aqueous Solutions of Electrolytes**

*The Electrical Conductivity of Aqueous Solutions at Elevated Temperatures and Pressures*

On the Calculation of the Conductivity of Aqueous Solutions Containing the Chlorides of Sodium and Barium

The Electrical Conductivity of Aqueous Solutions. A Report ... Upon a Series of Experimental Investigations Executed by A. A. Noyes, W. D. Coolidge, A. C. Melcher, Etc

**The Conductivity and Ionization of Electrolytes in Aqueous Solutions as Conditioned by Temperature, Dilution and Hydrolysis**

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The Electrical  
Conductivity of Aqueous  
Solutions of Sodium  
Chloride and Magnesium  
Sulphate at High  
Frequencies

**The Electrical  
Conductivity of Some  
Concentrated Aqueous  
Solutions**

The Electrical  
Conductivity of  
Aqueous Solutions  
of Sodium  
Monofleurophosphate  
and Potassium  
Hexafluorophosphate