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the rate of the forward reaction ( $k_f$ ) is equal to the rate of the reverse reaction ( $k_r$ ) in what is described as a dynamic equilibrium such that no change in their concentrations is observed. *Equilibrium A Reversible Reaction*

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Increase (T) causes the equilibrium to shift to the right (+) Decrease (I) causes the equilibrium to shift to the left Increase (T) causes the equilibrium to shift to the left Decrease (I) causes the equilibrium to shift to the right Temperature: A change in temperature corresponds to a change in energy therefore by using *CHEMICAL EQUILIBRIUM (ICE METHOD)*

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