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The eleventh edition was carefully reviewed with an eye toward strengthening the content available in OWLv2, end-of-chapter questions, and updating the presentation. Nomenclature changes and the adoption of IUPAC periodic table conventions are highlights of the narrative revisions, along with changes to the discussion of d orbitals. In-text examples have been reformatted to facilitate learning ...

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B) The lattice energy for a solid with 2+ and 2- ions should be two times that for a solid with 1+ and 1- ions. C) MgO has a larger lattice energy than LiF. D) Lattice energy is often defined as the change in energy that occurs when an ionic solid is separated into isolated ions in the gas phase.

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Darrell Ebbing has taught general chemistry for more than thirty years and is now retired from Wayne State University. He received his Ph.D. in physical chemistry from Indiana University. Steven D. Gammon is the Provost at the University of Maine Fort Kent.

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