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 $r_0 = 1.090 \times 10^{-10} \text{ m}$
 $r_0 = 1.090(2.74 \times 10^{-10} \text{ m}) = 2.99 \times 10^{-10} \text{ m}$
 For the Ne crystal, $r_0 = 2.74 \times 10^{-10} \text{ m}$ and $E_0 = 0.003121 \text{ eV}$. Therefore, $r_0 = 1.090(2.74 \times 10^{-10} \text{ m}) = 2.99 \times 10^{-10} \text{ m}$ for Ne.
 Calculate energy per atom at equilibrium:
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