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40.0 cm^3 of water to give a saturated

solution at 298K. If the solubility of the salt is

20.0 $mol\ dm^{-3}$ at 298K, what percentage of

the salt is left undissolved? {K= 39, Cl = 35.5,

O = 16}

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Solution expected to form acidic or basic oxide. (b)

(i) State two assumptions of the ... WAEC Chemistry

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(2) Determination of concentration of liquid solutes (stock solution) given the density, w/v, w/w), specific gravity, relative molecular mass, molar mass, and % purity. (3) Primary standard, secondary standard and standardized solution. WAEC Syllabus For Chemistry | 5.0 ENERGY AND ENERGY CHANGES (a) Energy changes in physical and isolated systems.

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A solution of 0.20 mole of NaBr and 0.20 mole of MgBr₂ in 2.0 dm³ of water is to be analysed. ...
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State whether R would be expected to form acidic or basic oxide. (b) (i) State two assumptions of the ...

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