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Assessment Solutions Class Date Section Quiz: Concentration of Solutions In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. 1. When preparing 500. ml, of a 1.35 M aqueous solution of NaCl, what should you do after adding the correct amount of solute to a large beaker? a.

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b. color of the solution. c. concentration of the solution. d. voltage difference between two electrodes in the meter ' s probe. \_\_\_\_ 15. When you conduct an acid-base titration, a. the pH of the solution must go up. b. the pH of the solution must go down. c. the pH of the solution must be 7.0 at the end point. d. the equivalence point must be ...

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