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Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$
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804 Chapter 25 Section 25.2 (continued) Half-Life Discuss Explain that, for each element, there exists only a small range of neutron-to-proton ratios that produce stable nuclei. If a nucleus does not reflect a

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10.3 Percent Composition and Chemical Formulas 10

Assuming salicylic acid $\text{C}_7\text{H}_6\text{O}_3$ as the limiting reagent calculate the theoretical yield. Molar mass of $\text{C}_7\text{H}_6\text{O}_3 = 132.1 \text{ g/mol}$. Molar mass of aspirin = 180.1 g/mol

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Introduction to Chemistry 7 1.1 FOCUS Objectives 1.1.1 Identify five traditional areas of ... Review, Interpreting Graphics ... 8 Chapter 1 Areas of Study Because the scope of chemistry is vast, chemists tend to focus on one area.

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1.1 Chemistry 1

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25.2 Nuclear Transformations 25

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Hg Trial 2 mass of flask+stopper= 83.39 grams Mass of flask + stopper +condensed vapor = 83.82 grams Temperature of boiling water = 99 degrees celsius Barometric ...

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