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Pu 0.03 Na 32.33 and then; Pu Na Pu Pu 0.03 Na 32.33 Pu Na Pu 1 1 f 0.108 and 2.61 N Na a , Na Na \* Navgdro 1 a , Na N Pu a , Pu M Na 1 Pu \* Navgdro a , Pu M Pu a , Pu 2.11b a , Na 8e 4b and finally k f 2.222(sup ercritical )  
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Convert P to a CFG,

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and then convert the CFG to a PDA, using the two constructions given in Section 6.3. The result is a one-state PDA equivalent to P.

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A glass of water is known to contain  $6.6 \times 10^{24}$  atoms of hydrogen. How many atoms of deuterium ( $^2\text{H}$ ) are present? Solution. According to Table 11.2 in Appendix II, the isotopic abundance of  $^2\text{H}$  is 0.015 a/o. The fraction of the hydrogen, which is  $^2\text{H}$ , is therefore  $1.5 \times 10^{-4}$  The

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