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6 Chemical Bonding

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the rates of effusion of gases at the same temperature and pressure are inversely proportional to the square roots of their

molar masses Rate of effusion of A / B = $\sqrt{\frac{M_b}{M_a}}$

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You can test your readiness to proceed by answering the Review Questions at the end of the chapter. This might also be a good

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Review Previous Concepts Gases CHAPTER 11 1. What is a mole? What information does the molar mass of a substance relate about the substance? 2. What is the kinetic-molecular theory of matter? Key Concepts LOUISIANA

STANDARDS LA.PS.29 Predict the properties of a gas based on gas laws. (e.g., temperature, pressure, volume). (PS-H-C7) 340 CHAPTER 11

Modern Chemistry Chapter 11 Review

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample 2.

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CHAPTER 6 REVIEW Chemical Bonding MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. a.

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