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Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16 Modern chemistry chapter 3 test b answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

Assessment Chapter Test A

Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

Modern Chemistry Chapter 3 Test Answer Key - Answers Fanatic

The Atoms: The Building Blocks of Matter chapter of this Holt McDougal Modern Chemistry Companion Course helps students learn the essential lessons associated with atoms.

Modern Chemistry Chapter 3 Test

CHAPTER 3 TEST Class Atoms: The Building Blocks of Matter MULTIPLE CHOICE On the line at the left of each statement, write the letter of the choice best completes the statement or answers the question. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of a. energy

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Assessment Chapter Test B

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Modern Chemistry Chapter 3 Test Answer Key

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. _____ 1.

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Modern Chemistry Chapter 3 Atoms: The Building Blocks of Matter law of conservation of mass- mass is neither created nor destroyed during ordinary chemical reactions or physical changes e.g. 20 g A + 20 g B 40 g AB

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3 Atoms: The Building Blocks of Matter - Ms. Sheehan's ...

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.