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____ 5. Mendeleev is credited with developing the first successful. a. periodic table. b. method for determining atomic number. c. test for radioactivity. d. use of X rays. ____ 6. Mendeleev did not always list elements in his periodic table in order of increasing atomic mass because he grouped together elements with similar. a. properties. c
...

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5 The Periodic Law
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Modern Chemistry 47 Chapter Test Name Class Date Chapter Test A, continued ____ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

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CHAPTER REVIEW To give students practice under more ... 60 minutes to answer all of the ques- tions in this Standardized Test Preparatio Answer the following items on a separate piece of paper. MULTIPLE CHOICE I. In the modern periodic table, elements are ...

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CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev’s original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

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Modern Periodic Law - As per the modern periodic law, the chemical and physical properties of elements are the periodic function of their atomic numbers Modern chemistry chapter 5 review the periodic law answer key. Click here for periodic table class 10.

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.
____ 1.

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. In his periodic table, Mendeleev did not list all of the elements in order

CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT ANSWER Answer the following questions in the space provided. 1. Identify the major assumption of the VSEPR theory, which is used to predict the shape of atoms. Pairs of valence electrons repel one another. 2. In water, two hydrogen atoms are bonded to one oxygen atom. Why isn’t water a linear ...

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Modern Chemistry 91 Chapter Test Name Class Date Chapter Test B, continued 27. Distinguish between ionic crystals and metallic crystals. PART V Write the answers to the following questions on the line to the left, and show your work in the space provided. The molar enthalpy of fusion for water is 6.008 kJ/mol. 28.