

Modern Chemistry Review Chemical Bonding Answers

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SECTION 4 SHORT ANSWER Answer the following questions in the space

provided. 1. _____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile.

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Join a group of classmates from your inorganic chemistry course. Each group should have 3 or 4 students. Arrange a time to meet for an hour or so. Prior to the meeting, read part I (pp. 1367-1391) as well as the summary at the end of Linus Pauling's article (J. Am. Chem. Soc. 1931, 53, 1367-1400) on the nature of the chemical bond.

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Modern Chemistry 47 Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1.

_____ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2.

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CHAPTER 6 REVIEW. Chemical Bonding. SECTION 6-2. SHORT ANSWER Answer the following questions in the space

provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule. 3.

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The modern chemical bond While still in college, a young chemist by the name of Linus Pauling familiarized himself with Lewis's work and began to consider how it might be interpreted within the context of the newly developed field of quantum mechanics.

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chemical bonding that results from the electrical attraction between large numbers of cations (metals) and anions (nonmetals) covalent bonding. chemical bonding that results from the sharing of electron pairs between 2 atoms. electronegativity.

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a covalent bond in which three pairs of electrons are shared between two atoms. multiple bonds. double and triple bonds are referred to as multiple bonds; double bonds in general have greater bond energy and are shorter than single bonds; triple bonds are even stronger and shorter. resonance.

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[chemical bonding | Definition and Examples | Britannica](#)

Oxtoby, David W., Nachtrieb, Norman H. Principles of Modern Chemistry. 3rd Ed. Saunders College Publishing, New York, 1996. Pauling, Linus, C. The Nature of the Chemical Bond and the Structure of Molecules and Crystals; An Introduction to Modern Structural Chemistry. 3rd Ed. Cornell University Press, Ithaca, 1960.

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6 Chemical Bonding. CHAPTER 6 REVIEW. Chemical Bonding. SECTION 2. SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond

approach each other, the electron-proton attraction is stronger than the electron-electron and proton-proton repulsions.

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chemical bonding resulting from the attraction between metal atoms and the surrounding sea of electrons. malleable the ability of a substance to be hammered or beaten into thin sheets.

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Chemical Bond A mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. **CHAPTER 6 REVIEW Chemical Bonding**

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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner

electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron. **Chemical Bonding | Chemistry | Visionlearning** Losing an electron always results in an increase in bond strength because the positively charged ion produced holds the electrons it has more strongly. The light removes an electron from a bonding orbital, increasing the bond order and bond strength.

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remains constant. 2. The electrons involved in the formation of a covalent bond are a. transferred from one atom to another. b. found only in the s orbitals. @valence electrons. d. in filled orbitals. 3. Bond energy is the energy a. absorbed as a molecule forms.