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1 a)The relative atomic mass of
an element is the weighted
average mass of the isotopes of
the element. (1) It is measured
on a scale on which a carbon-12
atom has a mass of exactly 12
units.(1). b)The total mass of
100 atoms = $(75 \times 35) + (25 \times 37) = 3550$ (1) The average mass
of 1 atom = $3550/100 = 35.5$ (1)

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