

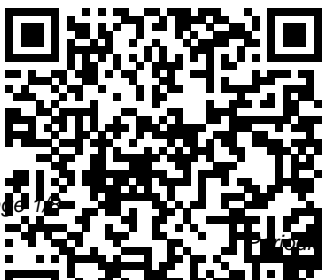
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# Periodic Table Section 3

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showing names,

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Answer: C. 3. An atom located in

the third period of the periodic

table would have an electron

configuration between

1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>1</sup> and

1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>. The highest

energy level is 3 as can be seen

from the electron configurations.

6.

100 Common

Questions About The

Periodic Table. A

Study Guide.

the periodic table

chapter 6 section 6.3

answer key, 6. How

many columns of

elements does the

---

periodic table contain?

\_\_\_18\_\_\_ 7. What is another name for a column of elements? \_\_\_group or family\_\_\_

8. What two group numbers can be used to designate elements in the second column of the periodic table?

\_\_\_group 2A or group 2\_\_\_ 9. How many rows of elements does the periodic table contain ...

The Periodic Table; Section 3 Flashcards | Quizlet

A section of the periodic table showing metals and non-metals.

The main groups are numbered from 1 to 7 going from left to right, and the last group on the right is group 0.

The section in the ... Section 3

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 arranged the elements by  
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