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Describe the following Periodic Trends, and give a theoretical explanation for each trend. a) atomic radius b) ionization energy c) electro negativity d) electron affinity How do we use theoretical explanation to describe this and in general Please and Thank You

Periodic trends of atomic radii? | Yahoo Answers

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Atomic Radius increases as you go towards the left and bottom of the periodic table, and Electron Affinity and Electronegativity increases as you go towards the right and up of the periodic table,...

[Periodic Trends: Electronegativity, Ionization Energy, Atomic Radius - TUTOR HOTLINE](#)

The Periodic Table: Atomic Radius, Ionization Energy, and Electronegativity Trends in the Periodic Table — Reactivity!

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SCIENCE 20-5 at C.d. Hylton High. Periodic Table Trends Practice 1. Rank the following elements by increasing atomic

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Periodic Trends Multiple Choice Review PSI Chemistry Name _____ Atomic Size 1) Elements Z and X are compared. Element Z is larger than Element X. Based on this you could say: A) Element Z is further to the

left side of the periodic table B) Element X is closer to the top of the periodic table

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1. Which of the following ions (Na^+ , Cl^- , Ca^{2+} , and/or S^{2+}) are isoelectronic with Ar? My Answer: Ca^{2+} , S^{2+} (WRONG) 2. Rank these elements (Kr, O, F) according to electron affinity by most energy released by gaining an electron or most energy absorbed by gaining an electron. My Answer: $\text{Kr} > \text{O} > \text{F}$ (WRONG) 3. Predict the ground-state electron configuration of the following ions. Write your ...

[what are the TRENDS in the periodic table? | Yahoo Answers](#)

The experiment we did was about Periodic Trends. We determined the trend of reactivity of nonmetals down a group, the trend of reactivity of metals along a period and the trend of acidic properties of elements along a period. The halogens used were Cl_2 , Br_2 and I_2 . Metals used were Sodium, Magnesium (magnesium ribbon), Aluminum. My questions: 1.)

Periodic Trends experiment question? | Yahoo Answers

the atomic number ALWAYS increases because of the number of protons in the nucleus. USUALLY the atomic mass increases (except for a few points in the periodic table such as iodine and tellurium)...

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Electronegativity follows a reverse trend to IE1. Within a period (as a given electron shell is filled), EN increases as the shell fills. So elements on the right of a period have greater EN than...

Periodic Trends (Rank these in order of... - Yahoo Answers

Periodic Trends: Electronegativity, Ionization Energy, Atomic Radius - TUTOR HOTLINE

The Periodic Table: Atomic Radius, Ionization Energy, and Electronegativity Trends in the Periodic Table — Reactivity!

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Ionization Energy - Basic Introduction

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1. Atomic size increases from top to bottom and decreases across the periodic table. 2. The ionization energy decreases from top to bottom and increases across the periodic table. 3. The...