
Practice 1. For a reaction where the rate equation is $r = k[\text{NH}_4^+(\text{aq})][\text{NO}_2^-(\text{aq})]$,
a) calculate k at temperature T_1 , if the rate, r , is $2.40 \times 10^{-7} \text{ mol}/(\text{L} \cdot \text{s})$ when $[\text{NH}_4^+$

RATE LAW EQUATION WORKSHEET 1

ANSWERS

Decreasing temperature (decreases, increases) the rate of reaction. 9. Enzymes are in molds and bacteria that spoil food. Explain, using your knowledge of factors affecting the rate of reaction, why food doesn't spoil as fast when it is refrigerated as it would at room temperature. 10.