

## Review Stoichiometry Section 1 Answers Modern Chemistry

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CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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CHAPTER 9 REVIEW Stoichiometry SECTION 9-3

PROBLEMS Write the answer on the line to the left.

Show all your work in the space provided. 1. 88% If the actual yield of a reaction is 22 g and the theoretical yield is 25 g, calculate the percent yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) → 2NH<sub>3</sub>(g) N<sub>2</sub>; 2.0 mol a.

**CHAPTER 8 REVIEW Chemical Equations and Reactions**

Stoichiometry Review Answers 1. a. Na<sub>3</sub>PO<sub>4</sub> b. Ca(NO<sub>3</sub>)<sub>2</sub> ...

Use the following balanced equation. Identify the limiting reactant when 1.150 grams of HgO react with 12.46 grams of Cl<sub>2</sub>. Convert to moles to get moles available, then calculate moles required:

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CHAPTER 8 REVIEW Chemical Equations and Reactions

Teacher Notes and Answers Chapter 8 SECTION 1

SHORT ANSWER 1. a. d b. a c. b d. f e. e f. c 2. 8,4,9 3.

a. 12 atoms ... SECTION 3 SHORT ANSWER 1. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg. 2. Fe forms an oxide in nature, and Ag does not, because it is much less active.

3. a. F<sub>2</sub> b. K c. H<sub>4</sub> a.

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REVIEW Stoichiometry SECTION 1 SHORT ANSWER

Answer the following questions in the space provided. 1.

\_\_\_\_\_ The coefficients in a chemical equation represent

the (a) masses in grams of all reactants and products. (b)

relative number of moles of reactants and products.

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New. CH 9 Mass to Mass Stoichiometry 2. CH 9 Limiting

Reactant . CH 9 Pract. Limit. Yield ... CH 11 Pre Test Answers.

CH 11 Section Review 11-1 - 11-3 . CH 11 Gas Stoichiometry

1 . CH 11 Gas Stoichiometry 2. CH 11 Boyles Law . CH 11

Charles Law . CH 11 Combined Gas Law-New ...

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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Section 9.1: Team Learning Worksheet 1. An individual

coefficient does not tell us anything. ... coefficient when doing stoichiometry calculations. For example, when given any mass of oxygen ... 1. The answer is "e" (each would produce the same amount of product). Many students try to answer this question without performing calculations ...

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11—1 Review and Reinforcement Stoichiometry ... stoichiometry

problem. A balanced equation verifies the law of If the statement is

true, write "true. " If it is false, change the underlined word or words

to make the statement true. Write your answer on the line provided

s. 10. 12. The term stoichiometr is derived from the words

**Chapter 9 Section 9.1: Team Learning Worksheet**

Stoichiometry. SECTION 1. SHORT ANSWER Answer the

following questions in the space provided. 1. \_\_\_\_\_ The

coefficients in a chemical equation represent the (a)

masses in grams of all reactants and products. (b) relative

number of moles of reactants and products. (c) number of

atoms of each element in each compound in a reaction.

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