

Solubility Aqueous Solutions

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[Chlorine Solubility in Aqueous Solutions](#) Hemisphere Pub

Excerpt from The Action of Water and Aqueous Solutions Upon Soil Carbonates The solubility in water of carbon dioxide, like all other gases, is greater at the lower temperatures than at the higher temperatures. With one or possibly two known exceptions, the solubility in aqueous solutions is decreased by increasing quantities of the material in solution. Thus, the solubility of carbon dioxide in water is decreased either by increasing the temperature or by the addition of some material, such as sodium chloride or other salts. The results of the work recorded in the literature have been assembled and are given in the following tables. The results are given in the same form as they have been recorded in the original papers. For instance, the solubility of carbon dioxide In water at 10° C. Has been given by Bunsen as This means that one cubic centimeter of water at 10° will dissolve the quantity of carbon dioxide occupying cubic centimeters at 0° and 760mm. All the gaseous volumes are reduced to 0° and 760 mm. Pressure. In this way comparisons may be made between the solubility of the gas in Solvents at different temperatures and also in different Solutions. About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books. Find more at www.forgottenbooks.com This book is a reproduction of an important historical work. Forgotten Books uses state-of-the-art technology to digitally reconstruct the work, preserving the original format whilst repairing imperfections present in the aged copy. In rare cases, an imperfection in the original, such as a blemish or missing page, may be replicated in our edition. We do, however, repair the vast majority of imperfections successfully; any imperfections that remain are intentionally left to preserve the state of such historical works.

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[Factors Affecting Solubility of Calcium Lactate in Aqueous Solutions](#)

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