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Unit 03: Solubility Equilibrium Multiple Choice Questions 1.

Which of the following would be true when equal volumes of 0.2 M NaBr and 0.2 M AgNO₃ are combined? A. No precipitate forms B. A precipitate of AgBr forms C. A precipitate of NaNO₃ forms D. Precipitates of both NaNO₃ and AgBr form 2.

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(a) Write the expression for the solubility-product constant, K_{sp}, and calculate its value at 18° C. (b) Calculate the equilibrium concentration of Mg²⁺ in 1.000 liter of saturated MgF₂ solution at 18°C to which 0.100 mole of solid KF has been added.

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Based on the solubility curves provided, both compound A and compound B will be completely in solution when dissolved in 100 g of water at 100 ° C. As the temperature of the solution is decreased to 0 ° C, 175 g of compound A will come out of the solution, while all of compound B will remain in the solution.

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Section (a) Write the expression for the solubility-product constant, K_{sp}, and calculate its value at 18 ° C. (b) Calculate the equilibrium concentration of Mg²⁺ in 1.000 liter of saturated

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Unit 3: Solubility - Ms. Bunney's Classes

A) The solubility increases B) The solubility decreases C) There is no

change in the solubility D) The change in the solubility is

unpredictable 36. Sugar water containing more sugar than would

normally dissolve at 20 degrees Celsius is said to be _____.

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