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The observed changes in the electrokinetic properties of calcite and apatite can be examined on the basis of the mineral-solution chemical equilibria involving dissolved mineral species. From studies of solubility isotherms for apatite and calcite at 25 ° C (Fig. 3.4, Fig. 3.5), the singular point for these minerals is identified to be 9.3. Above this pH calcite is more stable than apatite.

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ous systems involving salt-type minerals are examined in this work. Stability relations have been discussed on the basis of bulk equilibria.

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3. SOLUTION-MINERAL EQUILIBRIA PART 1: CARBONATES. Carbonic acid and the carbonate

minerals provide another good illustration of the use of equilibrium reasoning in geochemistry.

Interactions among these compounds determine the conditions under which limestones and dolomites are formed or dissolved, and likewise the conditions of formation of carbonate minerals as cements in soils and sandstones and as vein fillings.

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