

Stoichiometry Multiple Choice Test Id B Answers

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AP Chemistry: Stoichiometry – Multiple Choice Answers

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Stoichiometry MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) How many grams of hydrogen are in 46 g of C_2H_6 ? 1) A) 2.8 B) 184 C) 0.36 D) 1.5 E) 5.8 2) How many moles of carbon dioxide are there in 52.06 g of carbon dioxide? 2) A) 8.648 B) 0.84523 C) 1.34 D) 0.125 E) 0.250

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AP Chemistry: Stoichiometry – Multiple Choice Answers 44. What number of moles of O_2 is needed to produce 14.2 grams of P_4O_{10} from P ? (Molar Mass $\text{P}_4\text{O}_{10} = 284$) (A) 0.0500 mole (B) 0.0625 mole (C) 0.125 mole (D) 0.250 mole (E) 0.500 mole 4 P + 5 $\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$ >> 14.2 g P_4O_{10} x 1 mol P_4O_{10}

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Practice stoichiometry test Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. The coefficients in a chemical equation represent the a. masses, in grams, of all reactants and products. b. relative numbers of moles of reactants and products. c. number of atoms in each compound in a reaction.

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Consider the following balanced equation. $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 3\text{O}_2 \rightarrow 2\text{H}_3\text{C}_6\text{H}_5\text{O}_7 + 3\text{H}_2\text{O}$ Determine the mass of citric acid ($\text{H}_3\text{C}_6\text{H}_5\text{O}_7$) produced when 2.5 mol $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ is used.

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Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, respectively. The possible atomic weight of X is ... 6. c In multiple choice questions without a calculator, you must look for the "easy math" ? You will be most successful at ...

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stoichiometry multiple choice test id Practice stoichiometry test Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. The coefficients in a chemical equation represent the a. masses, in grams, of all reactants and products. b. relative numbers of moles of reactants and products. c.

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To understand stoichiometry, start with this introduction to the topic. It might also help to review molecules and moles, which includes how chemical formulas work. Ready for another quiz? Here's a quick self-test about the mole. If you'd rather switch gears, see if you know the answers about how chemistry explains the real world.

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Problem Eleven A 5.104 g sample of impure $\text{Na}_2\text{C}_2\text{O}_4$ was titrated with 30.55 mL of a 0.03928 M solution of NaMnO_4 , according to the equation: $2\text{NaMnO}_4 + 5\text{Na}_2\text{C}_2\text{O}_4 + 8\text{H}_2\text{SO}_4 \rightarrow 6\text{Na}_2\text{SO}_4 + 2\text{MnSO}_4 + 10\text{CO}_2 + 8\text{H}_2\text{O}$ What is the percentage of $\text{Na}_2\text{C}_2\text{O}_4$ in the sample?. a) 7.876% b) 4.523% c) 6.612%. Correct A look at the previous question will show that there is a 5 to 2 mole ...