
Stoichiometry Review Packet Answers

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Stoichiometry - Complete pages 1 and 2 in the practice packet (questions 1-7). The practice packet is linked above, with lots of instructions and an answer key. - This second video may help with larger problems with more steps. Pages 1 and 2 (questions 1-7) in the practice packet. Questions 1 and 2 were completed as examples in class.

Chapter 12 Stoichiometry Packet Answers

Chapter 12 Stoichiometry Packet Answer Key Answer: 4.93×10^{-5} L or $49.3 \mu\text{L}$ In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is

Stoichiometry Packet Answers Chapter 12

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W/ answers Website Upload 14. $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + \text{H}_2$ 15. $2 \text{ Al} + 3 \text{ HgCl}_2 \rightarrow 2 \text{ AlCl}_3 + 3 \text{ Hg}$ 16. $3 \text{ Ca(OH)}_2 + 2 \text{ H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6 \text{ H}_2\text{O}$ 17. $3 \text{ AgNO}_3 + \text{K}_3\text{PO}_4 \rightarrow \text{Ag}_3\text{PO}_4 + 3 \text{ KNO}_3$ 18. $\text{C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ Use the law of conservation of mass to determine the missing reactant in the equation given ...

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Packet. Tue 2/4. Practice Stoichiometry. Pgs 288-295.

Homework #3; Packet. Wed 2/5. Stoichiometry Part II Packet.

Thurs 2/6. Practice Stoichiometry Packet. Fri 2/7.

Review/Practice Stoichiometry Lab Calculations. Mon 2/10. No School Lincoln Day Review Notes. Tue 2/11. Limiting Reactants & Percent Yield. Pgs 296-307. Homework #4; Packet. Wed 2 ...

Unit 6: Reactions and Stoichiometry

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