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Chapter 11 –  
Stoichiometry Section  
11.1 What is  
stoichiometry? 1. true  
2. true 3. false 4. true  
5. true 6. 2, 2, 64.10 7.  
3, 3, 96.00 8. 2, 2,  
88.02 9. 4, 4, 72.08  
10. methanol and  
oxygen gas 11. carbon  
dioxide and water 12.  
160.10 g 13. 160.10 g  
14. They are equal.  
15. A mole ratio is a  
ratio between ...  
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a component of  
chemistry that  
analyzes the  
amount of  
substances that  
are involved in  
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Study Guide Chemistry Stoichiometry Answer The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

\_\_\_\_\_ 2.

Stoichiometry is based on the law of conservation of mass.

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*Answers* shows how to do  
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Stoichiometry; stoichiometry  
Chapter 10 - problems for  
States of Matter; which you are  
Chapter 11 - asked to convert  
Gases; ... Chapter from mass of one  
21 - Nuclear substance in a  
Chemistry; given chemical  
Chapter 22 - reaction to the  
Organic

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corresponding mass of another substance participating in the same reaction. For a related section, see Equation Stoichiometry Problems with Mixtures on our Web site.

## **Stoichiometry Study Guide**

### **Answers**

### **Chemistry**

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Answer The study

of the quantitative

relationships

between the

amounts of

reactants used

and the amounts

of products formed by a chemical reaction is called stoichiometry.

\_\_\_\_\_ 2. Stoichiometry is based on the law of conservation of mass. \_\_\_\_\_ 3.

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### **Answers**

Stoichiometry is a branch of chemistry and helps in the study of substances.

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Reviewing Vocabulary 1. In chemistry, these are called atoms - the basic units of an elemental form,

and molecules - basic units of compounds. 6 (467)  
Political Science. 2, the identity of the ...

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Stoichiometry Problems. When we carry out a reaction in either an industrial setting or a laboratory, it is easier to work with masses of substances than with the numbers of molecules or moles. We will first present the method used in most other books for converting from the mass of any reactant or product to the

mass of any other reactant or product using a balanced chemical equation outlined ...