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Stoichiometry Answer The study
of the quantitative relationships
between the amounts of reactants
used and the amounts of products
formed by a chemical reaction is
called stoichiometry. ____ 2.
Stoichiometry is based on the law
of conservation of mass.

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Calculations In your
textbook, read about mole-
to-mole conversion. Read
the following passage and
then solve the problems.
In the equation that
follows each problem,
write in the space
provided the mole ratio

that can be used to solve the problem.

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Stoichiometry Section

11.1 What is

stoichiometry? 1. true 2.

true 3. false 4. true 5. true

6. 2, 2, 64.10 7. 3, 3,

96.00 8. 2, 2, 88.02 9. 4,

4, 72.08 10. methanol and

oxygen gas 11. carbon

dioxide and water 12.

160.10 g 13. 160.10 g 14.

They are equal. 15. A

mole ratio is a ratio

between ...

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these are called atoms - the

basic units of an elemental

form, and molecules - basic

units of compounds. 6 (467)

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identity of the ...

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Stoichiometry is a branch

of chemistry and helps in

the study of substances.

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problems for which you

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a given chemical reaction

to the corresponding

mass of another substance participating in the same reaction. For a related section, see Equation Stoichiometry Problems with Mixtures on our Web site.

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Stoichiometry is based on the law of conservation of mass. _____ 3.